

1. **Capillary Action:** The attraction between water molecules and molecules of surrounding materials.
2. **Density of Water:** Density of water is 1 g/mL or 1 g/cubic centimeters. As water freezes into ice (solid) it expands and becomes less dense, about .93 g/mL, and will float in water.
3. **Freezing Point of Water:** Freezing is a change from liquid to solid(ice). Under normal circumstances, water will freeze or melt at 32 degrees Fahrenheit or 0 degrees Celsius.
4. **Hydrogen Bonding:** Molecules are attracted to each other.
 - This forms a strong bond.
 - Weaker than covalent bond of the water molecule
5. **Melting Point of Water:** Melting is the change of a solid (ice) to liquid. Under normal circumstances, water will freeze or melt at 32 degrees Fahrenheit or 0 degrees Celsius.
6. **Polarity:** When hydrogen and oxygen atoms bond the hydrogen end of the molecule becomes positively charged and the oxygen end becomes negatively charged.
7. **Solubility Of Water:** Water's polarity allows it to dissolve other substances easily. Water dissolves more substances than most other chemicals, so it is called the "universal solvent"
8. **Specific Heat Capacity of Water:** The specific heat of water is 1. Specific heat is the amount of heat energy needed to increase the temperature of 1 gram of water by 1 degree Celsius. Water has a high specific heat, so it takes longer to heat up and cool down. Remember: sea and land breezes from 7th grade.
9. **States of Matter:** Water is the only substance that naturally occurs as solid, liquid and gas in the atmosphere & on Earth's surface.
 - Gas: water vapor
 - Liquid: water
 - Solid: snow, glaciers
10. **Surface Tension:** Water has strong surface tension.
 - Molecules at the surface of water 'stick together' to form a 'skin' on the water.
 - Cohesion causes water to clump in droplets.
 - Adhesion causes water to stick to other polar substances.